

CLASS- 11 C

SUBJECT- CHEMISTRY

Things to remember

- Dalton's Atomic Theory.
- Laws of chemical combination.
- Avogadro's hypothesis.

WORKSHEET

FULL MARKS (20)

Q1

A) Fill in the blanks (4)

- 1) Formation of CO and CO₂ shows the law of _____.
- 2) The volume occupied by 23 grams of NO₂ at STP is _____ cc.
- 3) The number of atoms present in 32 gm of oxygen at STP is _____.
- 4) The number of electrons in 1.8 ml of water is _____.

B) Answer briefly (2*3)

- 1) Define the terms atomic mass and gram atomic mass.
- 2) Calculate the number of water molecules in 90 grams of water.
- 3) State the law of reciprocal proportions.

Question2 (2)

State and explain the law of multiple proportions.

Question 3 (2)

State the postulates of Dalton's Atomic Theory.

Question4 (2)

How much copper can be obtained from 100gm of CuSO₄? (Cu=63.5, S=32, O=16).

Question 5 (4)

State the following-

- 1) Avogadro's hypothesis
- 2) Gay Lusaac's law of gaseous volume

OR

State with reasons-

- 1) Why was it necessary to bring in the concept of molecules?
 - 2) What do you mean by mole concept?
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